IN THE SPECIFICATION

Please replace the paragraph beginning on page 11, line 1, with the following rewritten paragraph:

To a solution heated to 40°C obtained by dissolving t-butyl 1-chloro-2-fluorocyclopropane-1-carboxylate (tertiary butyl ester; cis/trans = 62/38; in this specification, regardless of the presence or absence of a halogen atom at the 1-position, a compound in which a fluorine atom and a carboxylic acid ester moiety are present on the same side of a plane of a cyclopropane ring is referred to as cis-form; 0.97 g, 5.0 mmoles) and tetrabutylammonium bromide (161 mg, 10% by mole) in methyl t-butyl ether (1.94 mL), an aqueous solution of sodium borohydride (concentration: 1 g/2.6 mL, 1.45 mL) was gradually added with stirring using a stirring blade. After the addition, the mixture was stirred using a stirring blade at 40°C for 20 hours, and then water was added to the reaction mixture. Diisopropyl ether was added to the mixture to conduct extraction (30 mL × 3) to give a diisopropyl ether solution containing 424 mg of the title compound (2a) (quantitation by high performance liquid chromatography, yield 53%). HPLC analysis conditions: column: MERCK Chromorith Performance RP-18 100-4.6 mm, mobile phase: pH 4.2 phosphate buffer/acetonitrile = 70/30, flow rate: 1.0 mL/min, detection wavelength: 220 nm. Further, gas chromatography analysis was conducted and cis/trans = 95/5 was found [analysis conditions: detector: ~~FDI~~ FID, column: GLscience, NEUTRA BOND-5, 30 m × 0.25 mm, temperature of vaporization chamber: 250°C, detector temperature: 250°C, carrier gas: nitrogen (80 kPa), hydrogen (60 kPa), air (50 kPa)].

Please replace the paragraph beginning on page 17, line 26, with the following rewritten paragraph:

By using the method of production of the present invention, the reaction time of a dehalogenation reaction of ~~2-halogeno-2-fluorocyclopropane-1-carboxylic acid ester~~ 1-halogeno-2-fluorocyclopropane-1-carboxylic acid ester can be significantly shortened as compared to the reaction time in previous methods. In particular, even in the case of using an apparatus for industrial production, the reaction can be completed in a shorter time. Furthermore, because dimethyl sulfoxide is not used as a reaction solvent in the method of production according to the present invention, the problem of generation of dimethyl sulfide has also been solved. Accordingly, the method of production of the present invention is industrially applicable as a method of producing a synthetic raw material for synthetic new quinolone antibacterial agents.